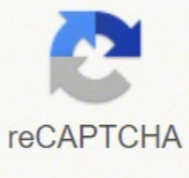




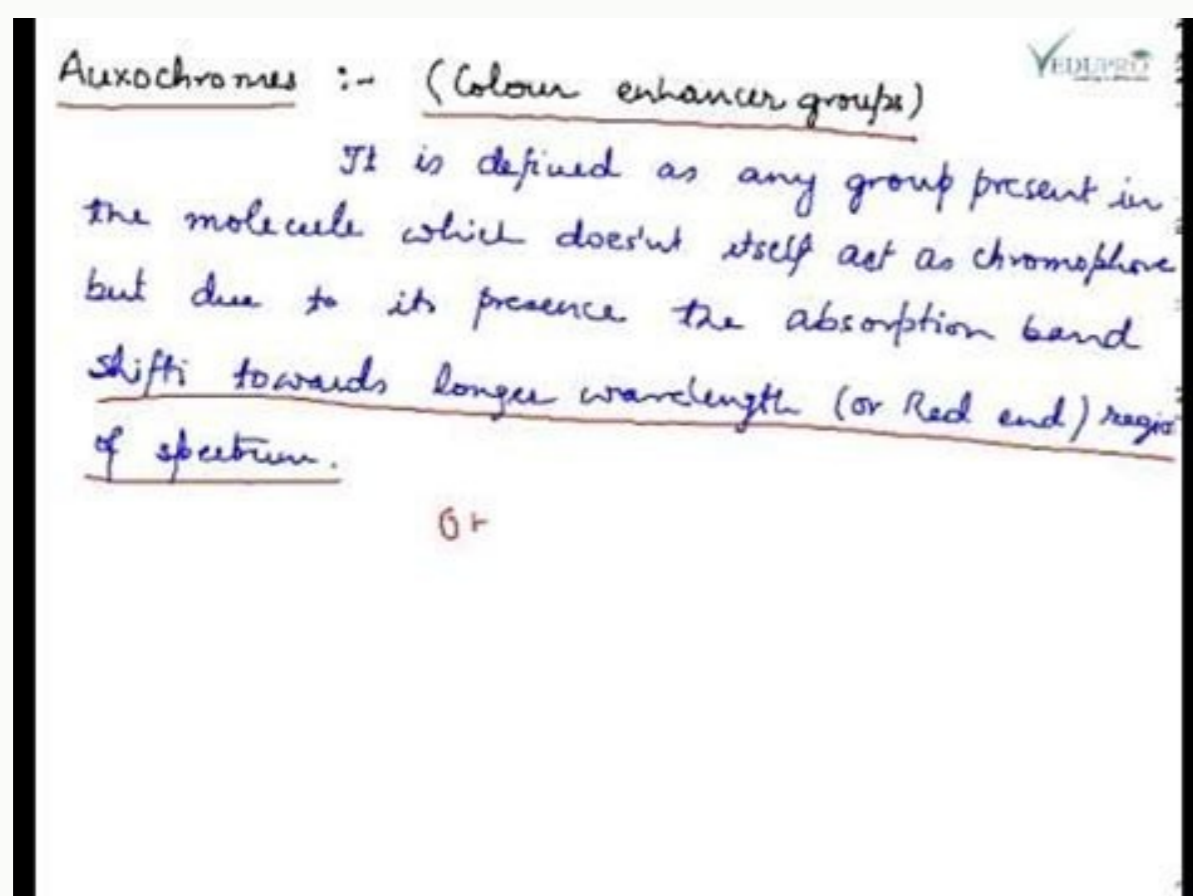
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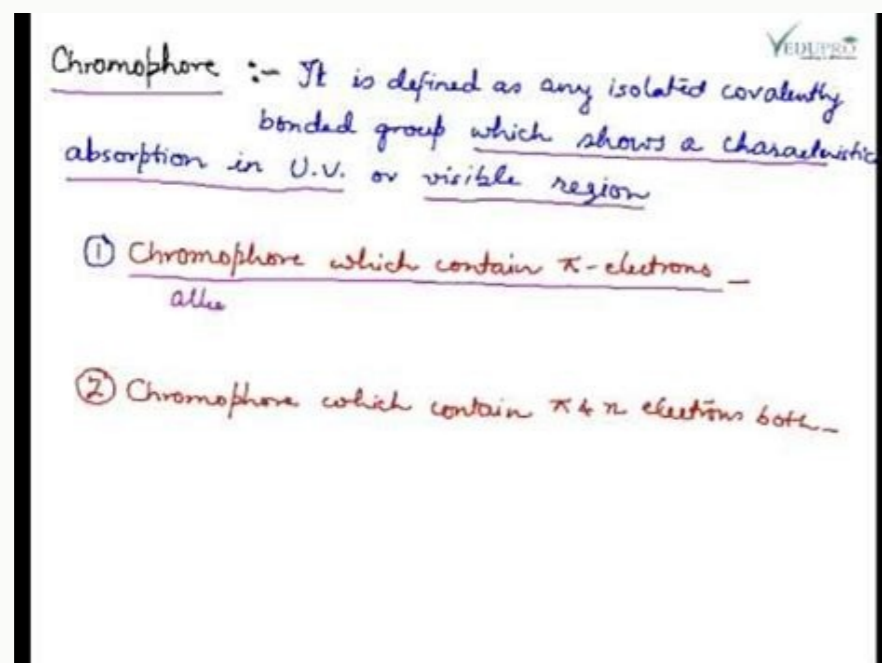
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What is chromophore and auxochrome give examples

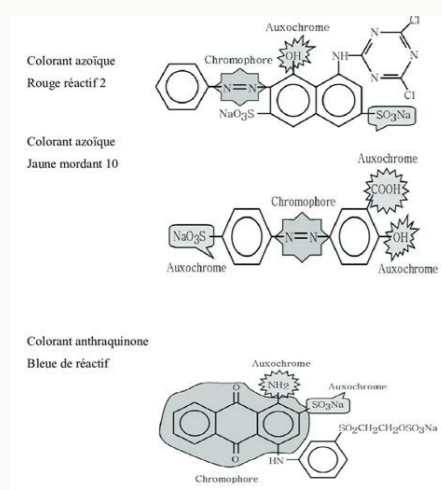
A functional group that changes a molecule's ability to absorb light. Additional quotes are required to verify this item. Please help improve this article by adding quotes to reliable sources. Unenriched equipment may be contested and discarded. Find sources: ASSOCHROME · NEWS · NEWSPAPERS · Books · sweetness · jstor (June 2015) (find out how and when to remove this message template) in organic chemistry, the auxochrome (from the ancient Greek augment $\pm A\frac{1}{2}i\eta\upsilon$ "e $\mu\acute{\iota}\mu\eta\iota$ (Chr re) "Color") is a group of atoms related to Achromofus. It modifies the ability of this chromophore to absorb light. They are unable to produce color on their own, but enhance the color of the chromogen when present with chromophores in an organic compound [1] Examples include hydroxyl (-oh), amino (-nh2), aldehyde (-cho), and methylmercapt ( sch3).[2] An auxochrome is a functional group of atoms with one or more free electron pairs that , when attached to a chromophore, change both the wavelength and intensity of absorption. When these groups are in direct conjugation with the PI system of the chromophore, they can increase the wavelength at which light is absorbed, resulting in absorption.The characteristic of these auxochromes is the presence of at least one free pair of electrons, which can be considered as a resonance magnification of the conjugated system. Effect on the chromophore. Increases the color of any organic compound. For example, benzene has no color because it has no chromophore; However, nitrobenzene is pale yellow because it contains a nitro group ( no2) that acts as a chromophore. However, p-hydroxynitrobenzene has a dark yellow color, the  OH group acting as an auxochrome. Here, the auxochrome ( oh) is combined with the chromophore  no2. Similar behavior is observed with azobenzene, which has a red color, but p-hydroxyazobenzene has a dark red color. The presence of goldThe B'Functional group that replaces the ability of the molecule to absorb light to test this article requires additional references. Please improve this article by involving references to reliable sources. The source can be challenged and removed. Find sources: Auxochrome \ XE2 \ X93 News \ XC2 \ XB7 Newspapers \ XC2 \ XC2 \ XC2 \ XC2 \ XB7 JStor (June 2015) Message Template (a \ XC5 \ X8D) \ "and increase \ " XCF \ X87 \ XCF \ X81 \ XE1 \ XB6 XCF \ XB1 (Chr \ XC5 \ X8DMA) \ "Color \) There is a group of atoms connected to the chrome. This changes the ability of this chromofor to absorb light. They are unable to produce color yourself, but they increase the color of the chromogen if they are accompanied by chromophs in an organic connection. [1] Examples are hydroxyl groups (\ xe2 \ x88 \ x92oh), aminograproups (\ xe2 \ x88 \ x92nh2), aldehyd (\ xe2 \ x88 \ x92cho) and methylkaptan (\ xe2 \ x88 \ x92oh) and methilkaptan (\ XE2 \ x88 \ x) (\ Xe2 \ x88 \ x) and methylk. that change both the wavelength and the absorption intensity in combination with chromophor. If these groups are in direct conjugation with the PI system chromophor, you can increase the wavelength in which the light is absorbed and thus absorbed. The main feature of these gold chrome is that there is at least one pair of electrons that can be regarded as a response in the conjugated system.



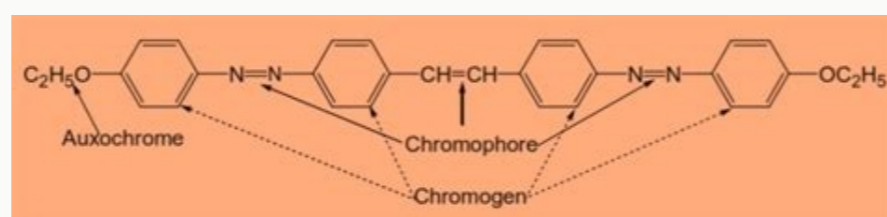
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Difference Between Auxochrome and Chromophore

Chromophore
porphyrin
porphyrin
heme
(chromophyll)
cytochromes
phytylaliphatic
carotenoids
chlorophyll
flavin
melanin

Auxochrome
M
O
P
N

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[ISBN is not] ^ Gronowitz, Salo J. O. (1958). Arkiv Fa. 27: 239. {{Cit Journal}: No or Blank | Name = (certificate) ^ "Dye language". New scientist. Information about the business industry. 122 (1665): 52. 1989 can ISSN 0262-4079. Source: " Check the answer using our 7 -day free test.These are: ethylene, acetylene, carbon, acids, foreign groups,

nitriles, etc. The carbon group is an important chromophore, although absorption of light by an isolated group does not produce color in ultraviolet spectroscopy. Types of chromophores: Two types of chromophores are known. 1. Chromophores that have electrons in their groups undergo changes. For example: -ylene, acetylene, etc. 2. Chromophores with electrons and N (non-bonding) electrons undergo two types of transitions. These are i-for" and n-for", for example: -Carbonyls, Nitriles, Azocomposited and nitrocomposited etc. Identification of chromophores: There are no hard and fast rules for identifying a chromophore. The change in position and intensity of absorption depends on many factors. The following points may be helpful. 1. A spectrum with a band of almost 300 Maken can have two or three units working together. 2. Absorption bands near 270-350 Macen with very low intensity Emax 10-100 are caused by n-absorption transitions of the Carbonico group, 3. Simple conjugated chromophores such as dienesor have maximum values of a α β δ -Curi α μ , i.e. 10000 to 20000. 4. An absorbance value between 1000 and 10000 indicates the presence of an aromatic system. When the aromatic core is replaced by groups capable of extending the chromophore, absorption occurs at values even higher than the extinction coefficients. Shift adsorption towards longer wavelengths as absorption intensity increases. Some commonly known auxochrome groups are -OH, -nh2, -or, -nhr and ânr2. For example: when the auxochrome group "NH2" is attached to the benzene ring. Its absorption ranges from "max 225" (Ēmax 203) to "max 280" (μ max1430). All auxochromes have one or more pairs of non-bonding electrons. When Auxocromo is attached to Chromo